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## Chapter 17 Thermochemistry Practice Problems Answers

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Chapter 1.

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Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 ° C.

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1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 ° C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21 ° C?

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Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 °C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21 °C? Explain your reasoning. 2.

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## Thermochemistry Guided Practice Problems

17 Practice Problems In your notebook, solve the following problems. SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK Use the three-step problem-solving approach you learned in Chapter many kilojoules of energy are in a donut that contains 200.0 Calories? 2. What is the specific heat of a substance that has a mass of 25.0 g and requires

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## Ch 17 Thermochemistry Practice Test Specific Heat Problems

1) How much heat must be absorbed by 375 grams of water to raise its temperature by  $25^{\circ}\text{C}$ ? 2) What mass of water can be heated from  $25.0^{\circ}\text{C}$  to  $50.0^{\circ}\text{C}$  by the addition of 2825 J? 3) What is the final temperature when 625 grams of water at  $75.0^{\circ}\text{C}$  loses  $7.96 \times 10^4\text{ J}$ ? Chapter 17.

## Enthalpy Problems And Solutions

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A 12.2 g sample of an unknown metal sample was  
heated to 98.6 ° C and then put into a calorimeter  
containing 25.0 ml of water at 22.3 ° C. After mixing,  
the temperature of the water and metal increased to a  
maximum of 28.4 ° C after 35 seconds.

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